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[EXPT 07 - THE IDEAL GAS LAW AND DENSITIES](#)

EXPERIMENT 7 THE IDEAL GAS LAW AND DENSITY In this experiment you will determine the average molecular mass of air using two different methods first by measuring the density of air with the density of several different gases and secondly by using the Ideal Gas Law To determine density of a gas you will fill a syringe with a gas and find its mass The volume of the gas will be determined by weighing a eq...

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[FORM 94 - SEARCH WARRANT \(NATIONAL ELECTRICITY LAW S.21 OR NATIONAL GAS LAW S.35\)](#)

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Ideal Gas Law Practice 1 What is the mass in grams of 125 mL of CO₂ at STP 0 246 grams of CO₂ 2 What is the density at STP of methane gas CH₄ in grams per liter 714 grams per liter 3 A chemical reaction produced 2 50 moles of nitrogen gas What volume in liters does this gas sample occupy at STP 56 liters 3 What is the molar mass of a gas that has a density of 0 840 g L at STP 18 8 g mol 4 A 0 215 g gas ...

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Dalton's Law of Partial Pressures $P_{total} = P_1 + P_2 + P_3 \dots$ Charles Law Boyle's Law $V_1 = V_2$ $V_1 P_1 = V_2 P_2$ T_1 T_2 Gay-Lussac's Law $P_1 = P_2$ T_1 T_2 Combined Gas Law Ideal Gas Law $V_1 P_1 = V_2 P_2$ $PV = nRT$ T_1 T_2 Some Pressure Conversion Factors 1 atmosphere = 760 mm Hg = 760 torr = 101.325 kPa Values for Standard Temperature and Pressure Pressure = 760 mm Hg Temperature = 273 K R Values .0821 atm L 62.4 mm ...

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