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NAME:

Name: Stoichiometry: Mole-Mole Problems
 1. $N_2 + 3H_2 \rightarrow 2NH_3$ How many moles of hydrogen are needed to completely react with 2.00 moles of nitrogen?
 2. $2KClO_3 \rightarrow 2KCl + 3O_2$ How many moles of oxygen are produced by the decomposition of 6.00 moles of potassium chlorate?
 3. $Zn + 2HCl \rightarrow ZnCl_2 + H_2$ How many moles of hydrogen are produced from the reaction of 3.00 moles of zinc with an excess of hydrochloric a...

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NAME:

Name: Classification of Chemical Reactions
 Directions: Classify the reactions below as synthesis, decomposition, single replacement(cationic or anionic) or double replacement.
 1. $2CO + O_2 \rightarrow 2CO_2$
 2. $2H_2O \rightarrow 2H_2 + O_2$
 3. $Ca(OH)_2 + H_2SO_4 \rightarrow CaSO_4 + 2H_2O$
 4. $2HgO \rightarrow 2Hg + O_2$
 5. $2H_2 + O_2 \rightarrow 2H_2O$
 6. $2KBr + Cl_2 \rightarrow 2KCl + Br_2$
 7. $CaO + H_2O \rightarrow Ca(OH)_2$
 8. $AgNO_3 + NaCl \rightarrow AgCl + NaNO_3$
 9. $2H_2O \rightarrow 2H_2 + O_2$
 10...

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NAME_

pH AND pOH CONTINUED Calculate the pH of the solutions below.
 1. 0.01 M HCl
 2. 0.0010 M NaOH
 3. 0.050 M $Ca(OH)_2$
 4. 0.030 M HBr
 5. 0.150 M KOH
 6. 2.0 M H_2CO_3 (Assume 5.0% dissociation.)
 7. 3.0 M HF (Assume 10.0% dissociation.)
 8. 0.50 M HNO_3
 9. 2.50 M NH_4OH (Assume 5.000% dissociation.)
 10. 5.0 M HNO_2 (Assume 1.0% dissociation.)
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