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AP Chemistry Lab #12 Lab #12: Determination of a Chemical Equilibrium Constant Objectives: 1. Determine the equilibrium constant of the formation of the thiocyanatoiron (III) ions. 2. Understand the application of using a spectrophotometer to measure concentration of colored ions. Materials: 5 Test Tubes (18 150 mm) 0.20 M iron (III) nitrate in 0.50 M (Flinn 0.60 M) of HNO<sub>3</sub> (aq) 4 Burets 0.0020 M iron (...)

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Chapter 13 - Chemical Equilibrium Intro A. Chemical Equilibrium 1. The state where the concentrations of all reactants and products remain constant with time 2. All reactions carried out in a closed vessel will reach equilibrium a. If little product is formed, equilibrium lies far to the left b. If little reactant remains, equilibrium lies far to the right 13.1 The Equilibrium Condition A. Static E...

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Chemical Equilibrium Consider a gas phase reaction of the form  $aA + bB \rightleftharpoons cC + dD$  We need to define a quantity known as the extent of reaction Then at any point in reaction time  $t$   $n_A = n_{A0} - \nu_A \xi$   $n_C = n_{C0} + \nu_C \xi$   $n_B = n_{B0} - \nu_B \xi$   $n_D = n_{D0} + \nu_D \xi$  Chemical Equilibrium As reaction proceeds will vary from 0 to some maximum value dictated by stoichiometry  $\xi_{max} = \min \left( \frac{n_{A0}}{\nu_A}, \frac{n_{B0}}{\nu_B}, \frac{n_{D0} + \nu_D \xi}{\nu_D} \right)$  In terms of free energy we can say  $G = G^0 + RT \ln Q$

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