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AP Chemistry Lab #12 Lab #12: Determination of a Chemical Equilibrium Constant Objectives: 1. Determine the equilibrium constant of the formation of the thiocyanatoiron (III) ions. 2. Understand the application of using a spectrophotometer to measure concentration of colored ions. Materials: 5 Test Tubes (18 150 mm) 0.20 M iron (III) nitrate in 0.50 M (Flinn 0.60 M) of HNO<sub>3</sub> (aq) 4 Burets 0.0020 M iron (...)

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Chapter 13 - Chemical Equilibrium Intro A. Chemical Equilibrium 1. The state where the concentrations of all reactants and products remain constant with time 2. All reactions carried out in a closed vessel will reach equilibrium a. If little product is formed, equilibrium lies far to the left b. If little reactant remains, equilibrium lies far to the right 13.1 The Equilibrium Condition A. Static E...

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Chemical Equilibrium Consider a gas phase reaction of the form  $aA + bB + cC + dD + eE + fF + gG$  We need to define a quantity known as the extent of reaction Then at any point in reaction time  $n_A, n_B, n_C, n_D, n_E, n_F, n_G$  Chemical Equilibrium As reaction proceeds will vary from 0 to some maximum value dictated by stoichiometry  $\Delta n_A, \Delta n_B, \Delta n_C, \Delta n_D, \Delta n_E, \Delta n_F, \Delta n_G$  Chemical Equilibrium In terms of free energy we can say  $G...$

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